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$P V / T = \text{constant}$ so $V_1 = (76 \times 1.0 \times 10^{-3} \times 290) / 273 \times 72 = 1.12 \times 10^{-3} \text{ m}^3$ 2. A mass of 1,200 cm³ of oxygen at 27°C and a pressure 1.2 atmosphere is compressed until its volume is 600 cm³ and its pressure is 3.0 atmosphere.

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